



HALF-LIFE PROBLEMS

Name Haley SmithBlock 2

1. An isotope of cesium (cesium-137) has a half-life of 30 years. If 1.0 g of cesium-137 disintegrates over a period of 90 years, how many g of cesium-137 would remain?

3 half-lives

$$\begin{array}{c} 1.0 \\ \downarrow \\ \frac{1}{2} \\ \downarrow \\ \frac{1}{4} \\ \downarrow \\ \frac{1}{8} \\ 1.25 \rightarrow 1.25 \end{array}$$

$$A = A_0 \left(\frac{1}{2}\right)^n$$

$$1.0 \text{ g} \left(\frac{\text{half-life}}{30 \text{ years}}\right) = 3 \text{ half-lives}$$

125 g

2. Actinium-226 has a half-life of 21.8 hours. If 100 mg of actinium-226 disintegrates over a period of 65 hours, how many mg of actinium-226 will remain?

$$5.8 \times \frac{1}{2} = 2 \text{ half-lives}$$

$$\begin{array}{c} 100 \\ \downarrow \\ \frac{1}{2} \\ \downarrow \\ \frac{1}{4} \\ \downarrow \\ \frac{1}{8} \\ 1.25 \end{array}$$

25 grams

3. Sodium-25 was to be used in an experiment, but it took 3.0 minutes to get the sodium from the reactor to the laboratory. If 0.00 mg of sodium-25 was removed from the reactor, how many mg of sodium-25 were placed in the reaction vessel 3.0 minutes later if the half-life of sodium-25 is 60 seconds?

3 1/2 half-lives

$$\begin{array}{c} 5 \\ \downarrow \\ \frac{1}{2} \\ \downarrow \\ \frac{1}{4} \\ \downarrow \\ \frac{1}{8} \\ 1.25 \rightarrow 1.25 \end{array}$$

62.5 g

4. The half-life of isotope X is 2.0 years. How many years would it take for a 4.00 mg sample of X to decay and have only 0.00 mg of it remain?

$$\begin{array}{c} 5 = 4 \left(\frac{1}{2}\right)^n \\ 125 = \left(\frac{1}{2}\right)^n \\ \log \log \\ 3 = n \end{array}$$

6 years

5. Selenium-83 has a half-life of 26.0 minutes. How many minutes would it take for a 10.0 mg sample to decay and have only 1.25 mg of it remain?

$$\begin{array}{c} 1.25 = 10 \left(\frac{1}{2}\right)^n \\ 10 = 10 \\ 12.5 = \left(\frac{1}{2}\right)^n \\ 2.5 = \frac{1}{2} \\ 2.5 = \frac{1}{2} \\ 2.5 = \frac{1}{2} \\ 2.5 = \frac{1}{2} \end{array}$$

3 1/2 half-lives

75.0 minutes

6. The half-life of Po-218 is three minutes. How much of a 2.0 gram sample remains after 15 minutes? Suppose you wanted to buy some of this isotope, and it required half an hour for it reach you. How much should you order if you need to use 0.10 grams of this material?

$$\begin{array}{c} 2.0 \\ \downarrow \\ 1.25 \\ \downarrow \\ 0.625 \end{array}$$

6 1/2 half-lives

$$\begin{array}{c} 2.0 \left(\frac{1}{2}\right)^6 \\ 2.0 \left(\frac{1}{2}\right)^6 \\ = 0.03125 \end{array}$$

$$\begin{array}{c} 10 = A_0 \left(\frac{1}{2}\right)^n \\ 10 = A_0 \left(\frac{1}{2}\right)^n \\ 10 = A_0 \left(\frac{1}{2}\right)^n \\ 10 = A_0 \left(\frac{1}{2}\right)^n \end{array}$$

60.0 grams

24 Practice Problems Half Life Answers

Michael R. Lindeburg

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