

Here are some practice problems for the current material dealing with molar masses. As an extra challenge, could anyone identify the substances below?

Calculate the formula masses for the compounds below.

- |                                      |   |  |
|--------------------------------------|---|--|
| a) $\text{H}_3\text{PO}_4$<br>98 amu | b) $\text{C}_4\text{H}_{10}$<br>58 amu                  | c) $\text{Co}(\text{NO}_3)_2$<br>183 amu   |
| d) $\text{FeCl}_3$<br>161 amu        | e) $\text{C}_{12}\text{H}_{22}\text{O}_{11}$<br>342 amu | f) $\text{Al}_2(\text{SO}_4)_3$<br>342 amu |

Calculate molar masses for the compounds below.

- |  |   |   |
|--|---|---|
| a) $\text{H}_2\text{O}$<br>18 g/mol                        | b) $\text{H}_2\text{SO}_4$<br>98 g/mol                                      | c) $\text{C}_6\text{H}_{12}\text{O}_6$<br>180 g/mol                       |
| d) $\text{C}_7\text{H}_8\text{N}_4\text{O}_2$<br>180 g/mol | e) $(\text{Na})\text{C}_{11}\text{H}_{18}\text{N}_2\text{O}_3$<br>249 g/mol | f) $(\text{C}_5\text{H}_5\text{N})_4\text{Cu}(\text{NCS})_6$<br>728 g/mol |

Calculate the mass of the following compounds.

- |                           |  |   |
|---------------------------|--|---|
| a) 1.00 mol NaCl<br>58 g  | b) 2.00 mol $\text{MgBr}_2$<br>368 g             | c) 0.50 mol $\text{CCl}_4$<br>76 g      |
| d) 3.60 mol NaOH<br>144 g | e) 0.710 mol $\text{Ca}(\text{CN})_2$<br>65.32 g | f) 1.82 mol $\text{BaSO}_4$<br>424.06 g |

Calculate the number of moles of the following compounds.

- |                                     |   |  |
|-------------------------------------|---|--|
| a) 88 g $\text{CO}_2$<br>2.00 moles | b) 198 g $\text{C}_2\text{H}_{12}\text{O}_6$<br>1.5 moles | c) 65.25 g LiBr<br>0.75 moles                        |
| a) 249g NaOH<br>6.225 moles         | b) 878g $\text{NaAlSi}_3\text{O}_8$<br>3.35 moles         | c) 7,327g $\text{Cu}_2\text{FeSnS}_4$<br>17.00 moles |

# 24 Practice Problems Chemistry

**G Psacharopoulos**



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