



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Practice Calculations

Use the following equations to calculate the pH of a solution with a given concentration of hydrogen ions.

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions \times concentration of other ions

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions \times concentration of other ions

Answers

1. $\text{pH} = -\log_{10}[\text{H}^+]$	2. $\text{pH} = \log_{10}[\text{OH}^-]$
3. $\text{pH} = -\log_{10}[\text{H}^+]$	4. $\text{pH} = \log_{10}[\text{OH}^-]$
5. $\text{pH} = -\log_{10}[\text{H}^+]$	6. $\text{pH} = \log_{10}[\text{OH}^-]$
7. $\text{pH} = -\log_{10}[\text{H}^+]$	8. $\text{pH} = \log_{10}[\text{OH}^-]$
9. $\text{pH} = -\log_{10}[\text{H}^+]$	10. $\text{pH} = \log_{10}[\text{OH}^-]$
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15. $\text{pH} = -\log_{10}[\text{H}^+]$	16. $\text{pH} = \log_{10}[\text{OH}^-]$
17. $\text{pH} = -\log_{10}[\text{H}^+]$	18. $\text{pH} = \log_{10}[\text{OH}^-]$
19. $\text{pH} = -\log_{10}[\text{H}^+]$	20. $\text{pH} = \log_{10}[\text{OH}^-]$
21. $\text{pH} = -\log_{10}[\text{H}^+]$	22. $\text{pH} = \log_{10}[\text{OH}^-]$
23. $\text{pH} = -\log_{10}[\text{H}^+]$	24. $\text{pH} = \log_{10}[\text{OH}^-]$
25. $\text{pH} = -\log_{10}[\text{H}^+]$	26. $\text{pH} = \log_{10}[\text{OH}^-]$
27. $\text{pH} = -\log_{10}[\text{H}^+]$	28. $\text{pH} = \log_{10}[\text{OH}^-]$
29. $\text{pH} = -\log_{10}[\text{H}^+]$	30. $\text{pH} = \log_{10}[\text{OH}^-]$
31. $\text{pH} = -\log_{10}[\text{H}^+]$	32. $\text{pH} = \log_{10}[\text{OH}^-]$
33. $\text{pH} = -\log_{10}[\text{H}^+]$	34. $\text{pH} = \log_{10}[\text{OH}^-]$
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