

12.3**LIMITING REAGENT AND PERCENT YIELD****Section Review****Objectives**

- Identify and use the limiting reagent in a reaction to calculate the maximum amount of product(s) produced and the amount of excess reagent
- Calculate theoretical yield, actual yield, or percent yield given the appropriate information

Vocabulary

- limiting reagent
- excess reagent
- theoretical yield
- actual yield
- percent yield

Key Equations

- $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Whenever quantities of two or more reactants are given in a stoichiometric problem, you must identify the 1. This is the reagent that is completely 2 in the reaction. The amount of limiting reagent determines the amount of 3 that is formed.

When an equation is used to calculate the amount of product that will form during a reaction, the value obtained is the 4. This is the 5 amount of product that could be formed from a given amount of reactant. The amount of product that forms when the reaction is carried out in the laboratory is called the 6.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 7. Normally, the actual yield in a chemical reaction will be equal to or less than the theoretical yield.
- _____ 8. The actual yield of a chemical reaction can be calculated using mole ratios.

12 3 Limiting Reagent And Percent Yield Answer Key

**John W. Moore, Conrad L.
Stanitski, Peter C. Jurs**



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