

## CHEMISTRY 103 – Practice Sheet #4 – Naming Compounds and Writing Formulas

## Chapter 2

Prepared by Dr. Tony Jacob

<http://www.chem3tsc.edu/nastaiclc> (Resource page)

1. Name the compounds below. Includes molecular and ionic compounds, polyatomic ions, acids, and transition metals.

a.  $\text{Na}_2\text{S}$  sodium sulfideb.  $\text{CaO}$  calcium oxidec.  $\text{N}_2\text{O}$  dinitrogen monoxided.  $\text{K}_3\text{N}$  potassium nitridee.  $\text{NO}$  nitrogen oxidef.  $\text{Li}_3\text{P}$  lithium phosphideg.  $\text{NaNO}_3$  sodium nitrateh.  $\text{Ca}_3(\text{PO}_4)_2$  calcium phosphatei.  $\text{MgF}_2$  magnesium fluoridej.  $(\text{NH}_4)_2\text{SO}_4$  ammonium sulfatek.  $\text{P}_2\text{O}_9$  diphosphorus pentoxidel.  $\text{NaOH}$  sodium hydroxidem.  $\text{HBr}$  hydrogen bromiden.  $\text{Sc}(\text{OH})_3$  scandium III hydroxideo.  $\text{HNO}_3$  nitric acidp.  $\text{NH}_4\text{NO}_3$  ammonium nitrateq.  $\text{FeCl}_3$  ferric chlorider.  $(\text{NH}_4)_2\text{O}$  ammonium oxides.  $\text{CuO}$  copper oxidet.  $\text{HCl}$  hydrochloric acidu.  $\text{NaHCO}_3$  sodium bicarbonatev.  $\text{Al}(\text{C}_2\text{H}_5\text{O}_2)_3$  aluminum acetatew.  $\text{H}_2\text{SO}_4$  dihydrogen sulfatex.  $\text{HI}$  hydroiodic acidy.  $\text{SO}_3$  sulfur trioxidez.  $\text{HC}_2\text{H}_3\text{O}_2$  acetic acidaa.  $\text{Ca}(\text{OH})_2$  calcium hydroxidebb.  $\text{Mn}(\text{ClO}_4)_3$  manganese III perchloratecc.  $\text{HClO}_4$  perchlorate aciddd.  $\text{Zn}(\text{CN})_2$  zinc II cyanideee.  $\text{NaC}_2\text{H}_3\text{O}_2$  sodium acetateff.  $\text{K}_2\text{SO}_4$  potassium sulfategg.  $\text{H}_2\text{CO}_3$  dihydrogen carbonatehh.  $\text{H}_3\text{PO}_4$  trihydrogen phosphateii.  $\text{HF}$  hydrofluoric acidjj.  $\text{CH}_3\text{COOH}$  acetic acidkk.  $\text{LiOH}$  lithium hydroxidell.  $\text{H}_3\text{O}^+$  trihydroxide ionmm.  $\text{NH}_3$  ammoniann.  $\text{CH}_4$  methane

# 103 Chemistry Worksheet Answers

**Terry C. Jones**

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### **Table of Contents 103 Chemistry Worksheet Answers**

1. Understanding the eBook 103 Chemistry Worksheet Answers
  - The Rise of Digital Reading 103 Chemistry Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 103 Chemistry Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 103 Chemistry Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 103 Chemistry Worksheet Answers
  - Personalized Recommendations
  - 103 Chemistry Worksheet Answers User Reviews and Ratings
  - 103 Chemistry Worksheet Answers and Bestseller Lists

5. Accessing 103 Chemistry Worksheet Answers Free and Paid eBooks
  - 103 Chemistry Worksheet Answers Public Domain eBooks
  - 103 Chemistry Worksheet Answers eBook Subscription Services
  - 103 Chemistry Worksheet Answers Budget-Friendly Options
6. Navigating 103 Chemistry Worksheet Answers eBook Formats
  - ePUB, PDF, MOBI, and More
  - 103 Chemistry Worksheet Answers Compatibility with Devices
  - 103 Chemistry Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 103 Chemistry Worksheet Answers
  - Highlighting and Note-Taking 103 Chemistry Worksheet Answers
  - Interactive Elements 103 Chemistry Worksheet Answers
8. Staying Engaged with 103 Chemistry Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 103 Chemistry Worksheet Answers
9. Balancing eBooks and Physical Books 103 Chemistry Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 103 Chemistry Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 103 Chemistry Worksheet Answers
  - Setting Reading Goals 103 Chemistry Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 103 Chemistry Worksheet Answers
  - Fact-Checking eBook Content of 103 Chemistry Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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