

Name _____

KEY

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REVIEW QUESTIONS FOR FINAL EXAM (2nd Semester)

SECTION 12.1

Heat Transfer: $Q = mC\Delta T$ or $Q = mc(C_f - C_i)$

Specific Heat of Common Substances			
Material	Specific Heat $J/kg \cdot K$	Material	Specific Heat $J/kg \cdot K$
aluminum	896	lead	126
iron	450	mercury	2350
carbon	714	silver	236
copper	386	water	4200
glass	800	ice	2090
oil	2090	iron	450
water	4200	lead	126

1) Two solid objects are in contact as shown below.

Heat will always flow from object A to object B if object B has a lower
a) mass b) total energy c) specific heat
d) temperature

For questions 2-3: The diagram below shows two solids with their temperatures ($A=50^\circ\text{C}$, $B=80^\circ\text{C}$).

2) What occurs when the two solids are placed in contact with each other as shown?

a) Heat energy flows from A to B. Solid A cools down.
b) Heat energy flows from A to B. Solid A gets hotter.
c) Heat energy flows from B to A. Solid B cools down.
d) Heat energy flows from B to A. Solid B gets hotter.

3) The net flow of heat energy will stop from one object to the other when they both have the same

a) energy b) temperature
c) specific heat d) all the answers are correct

4) How much heat is transferred when a 0.2 kg piece of copper's temperature is raised from 50°C to 80°C ?

$$Q = mc\Delta T$$

$$Q = (0.2\text{ kg})(385\text{ J/kg}\cdot\text{C})(80^\circ\text{C} - 50^\circ\text{C}) = 2310\text{ J}$$

5) A 0.5 kg block of lead cools down from 30°C to 20°C . How much heat was transferred?

$$Q = (0.5\text{ kg})(130\text{ J/kg}\cdot\text{C})(20^\circ\text{C} - 30^\circ\text{C})$$

$$Q = -650\text{ J}$$

6) The heat transfer of a substance is 9400 J. The mass of this substance is 2 kg and its temperature rises from 80°C to 100°C . Based on the specific heat that you will calculate for this substance, what is it?

$$C = \frac{Q}{m\Delta T} = \frac{9400\text{ J}}{(2\text{ kg})(100^\circ\text{C} - 80^\circ\text{C})} = 235\text{ J/kg}\cdot\text{C}$$

SILVER

7) From the choices below, which material would be better for a pan handle?

a) brass 376 b) silver 235 c) copper 385
d) iron 450 e) zinc 377 f) lead 120

8) Provide a brief description of absolute zero and their temperatures in Kelvin (K) and degrees Celsius ($^\circ\text{C}$).

The lowest temperature when all particle motion stops. $T_c = -273^\circ\text{C}$, $T_k = 0\text{ K}$.

9) The boiling point of water is 100°C . The boiling point of water in Kelvin (K) will be

$$T_k = 273 + T_c = 273 + 100 = 373\text{ K}$$

10) A 1-liter of water taken out of a 5-gallon bucket full of water at 20°C will

a) register a lower temperature than that of the water left in the bucket.
b) have less thermal energy than that of the water left in the bucket.

c) have less specific heat than that of the water left in the bucket.

11) 243 K in degrees Celsius ($^\circ\text{C}$) is

$$T_c = T_k - 273 = 243\text{ K} - 273 = -30^\circ\text{C}$$

12) The kinetic molecular theory can explain the existence of thermal energy in a substance.

a) true b) false

13) When using a thermometer for measuring the temperature of a substance, one should record the temperature

a) about 2 minutes after the thermometer has been placed in the substance. That is usually the time needed for the liquid inside to expand or shrink completely.

b) until thermal equilibrium has been reached

c) until all the heat from the substance has been transferred to the thermometer

d) until the thermal energy from the substance is equal to the thermal energy of the thermometer

e) all of the answers

14) A calorimeter contains 1 kg of water at 20°C . A 0.20 kg block of zinc at 120°C is placed in the water. What is the approximate final temperature of the system?

$$T_f = \frac{m_A C_A T_A + m_B C_B T_B}{m_A C_A + m_B C_B}$$

$$T_f = \frac{(1)(4110)(20^\circ\text{C}) + (0.2)(387)(120^\circ\text{C})}{(1)(4110) + (0.2)(387)} = 21.8^\circ\text{C}$$

$$T_f = \frac{8360 + 9312}{4257.6} = 21.8^\circ\text{C}$$

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Lei Huang

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