

CHAPTER 5: ATOMIC STRUCTURE

Atom consists of several subatomic particles. The main parts of an atom are:

Nucleus: central part of an atom, consists of two types of particles

- Protons:** positive charged particles. Has a relative mass of 1 atomic mass unit (u)
- Neutrons:** neutral particles. Has a relative mass of 1 atomic mass unit (u)

Electron Cloud: surrounding the nucleus, a cloud-like region divided into energy levels or electron shells.

- Electron – negatively charged particle**.

Subatomic particles – determines the properties of an atom, including its chemical behavior and reactivity

Number of protons in nucleus defines the element's identity.

Number of neutrons can vary within the same element, resulting in isotopes.

Number of electrons determines the atom's overall charge.

Number of protons = number of electrons

Electrons in Energy Levels and Sublevels

Principal Energy Level (n)	Sublevels available	# of orbitals in sublevel (n ²)	# of electrons possible in sublevel	Total electrons for energy level (2n ²)
1	s	1	2	2
2	s p	1 3	2 6	8
3	p d	1 3 5	2 6 10	18
4	p d f	1 3 5 7	2 6 10 14	32

Isotopes

- Have different mass number but same atomic number.
- are members of a family of an element that have the same number of proton but different number of neutron.

How to Get Neutron?

$$n = A - Z$$

A = mass number

Z = atomic number

How to Get Atomic Mass?

$$A = Z + n$$

Z = atomic number

n = neutron

Ion – is an element that has different protons and electrons.

Cation – positively charged ion; more protons than electrons.

Anion – negatively charged ion; more electrons than protons.

REPORTING

Relative Mass – means that if the electron has a mass of 1 unit, the proton and neutron will have a mass 2000 times that of the electron.

Proton – discovered by Eugene Goldstein

Electron – discovered by Joseph John Thompson

Neutron – discovered by Sir James Chadwick

Isotopes – atoms of the same element that have different mass (amu)

Angstrom – a common unit for describing the sizes of atoms

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James N. Jensen

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