

# Acids & Bases HOMEWORK

## Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO<sub>3</sub>**: **nitric acid**
- d. **HClO<sub>4</sub>**: **perchloric acid**
- e. **HBrO<sub>3</sub>**: **hypobromous acid**
- f. **HClO<sub>4</sub>**: **chloric acid**
- g. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- h. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- i. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**
- j. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**

2. Of the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. "Write a balanced equation for the reaction:  $\text{Ba(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + \text{H}_2\text{O}$ "

4.  $\text{HClO}_4$  will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_3\text{O}^+ + \text{HClO}_4 \rightarrow \text{H}_2\text{O} + \text{HClO}_4$

b. Write a balanced equation for a small equilibrium constant for this reaction? **weakly** because  $\text{HClO}_4$  is a strong acid

c. Write a balanced equation for a small equilibrium constant for this reaction? **strongly** because  $\text{H}_3\text{O}^+$  is a strong acid

d. In your reaction equation, is  $\text{HClO}_4$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a proton

e. Identify the conjugate base of  $\text{HClO}_4$ :  **$\text{ClO}_4^-$**

f. Identify  $\text{HClO}_4$  as aprotic, amphotropic, or polyprotic. **polyprotic** because it can lose multiple  $\text{H}^+$  ions

5.  $\text{H}^+$  will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_3\text{O}^+ + \text{H}_2\text{O} \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$

b. Write a balanced equation for a small equilibrium constant for this reaction? **weakly** because  $\text{H}_3\text{O}^+$  is a strong acid

c. Write a balanced equation for a small equilibrium constant for this reaction? **strongly** because  $\text{H}_2\text{O}$  is a strong base

d. In your reaction equation, is  $\text{H}^+$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a proton

e. Identify the conjugate base of  $\text{H}^+$ :  **$\text{H}_2\text{O}$**

f. Identify  $\text{H}^+$  as aprotic, amphotropic, or polyprotic. **aprotic**

6.  $\text{HClO}_4$  will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_3\text{O}^+ + \text{HClO}_4 \rightarrow \text{H}_2\text{O} + \text{HClO}_4$

b. Write a balanced equation for a small equilibrium constant for this reaction? **weakly** because  $\text{HClO}_4$  is a strong acid

c. Write a balanced equation for a small equilibrium constant for this reaction? **strongly** because  $\text{H}_3\text{O}^+$  is a strong acid

d. In your reaction equation, is  $\text{HClO}_4$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a proton

e. Identify the conjugate base of  $\text{HClO}_4$ :  **$\text{ClO}_4^-$**

f. Identify  $\text{HClO}_4$  as aprotic, amphotropic, or polyprotic. **polyprotic** because it can lose multiple  $\text{H}^+$  ions in the reaction (not like you said)

## 19 3 Acids And Bases Answer Key Chemistry

**Thomas Edward Thorpe**

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## **Table of Contents 19 3 Acids And Bases Answer Key Chemistry**

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
  - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry
  - Personalized Recommendations
  - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
  - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists
5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks

- 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
- 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
- 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options

6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats

- ePUB, PDF, MOBI, and More
- 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
- 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
- Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
- Interactive Elements 19 3 Acids And Bases Answer Key Chemistry

8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry

9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry

- Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry

- Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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synapses brain and reflexes notes with practice questions at the end 7

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