

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given pK_a or pK_b
3. Calculate the percent ionization of weak acids

Vocabulary

- | | | |
|----------------------------|-------------------|-----------------------------|
| acid dissociation constant | 1. conjugate acid | 2. weak acid |
| acid ionization | 3. base | 4. base ionization constant |
| | 5. weak base | |

Key Equations

- | | |
|------------------------------------|-----------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[HB][OH^-]}{[B]}$ |
| 3. $pK_a = -\log K_a$ | 4. $pK_b = -\log K_b$ |

Part A Completion

Write the appropriate number in each blank space, or write the appropriate chemical formula in each blank space, or complete each sentence.

1. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
2. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
3. The pK_a of a weak acid is 4. The pK_b of its conjugate base is _____.
4. The pK_b of a weak base is 4. The pK_a of its conjugate acid is _____.
5. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
6. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
7. The pK_a of a weak acid is 4. The pK_b of its conjugate base is _____.
8. The pK_b of a weak base is 4. The pK_a of its conjugate acid is _____.
9. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
10. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.

1. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
2. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
3. The pK_a of a weak acid is 4. The pK_b of its conjugate base is _____.
4. The pK_b of a weak base is 4. The pK_a of its conjugate acid is _____.
5. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
6. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
7. The pK_a of a weak acid is 4. The pK_b of its conjugate base is _____.
8. The pK_b of a weak base is 4. The pK_a of its conjugate acid is _____.
9. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
10. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.

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