

KEY

AP CHEMISTRY MULTIPLE CHOICE QUESTIONS
Unit 8 – Introduction to Equilibrium

The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



(A) 95

(B) 95

(C) 95^2 (D) $\frac{1}{95}$ (E) $\frac{1}{(95)^2}$

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products? *increase P*



Shift towards fewer gas

4.



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

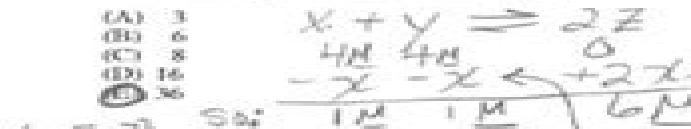
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6.

The synthesis of $CH_3OH(g)$ from $CO(g)$ and $H_2O(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.

(B) K_c will increase because there are more reactant molecules than product molecules.

(C) K_c will decrease because the reaction is exothermic.

(D) K_c is constant and will not change.



$C_2H_6(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- Changes in the partial pressure of $H_2(g)$
- Changes in the particle size of the platinum catalyst
- Changes in the temperature of the reaction system

(A) III only
(B) I and II only
(C) I and III only
(D) II and III only
(E) I, II, and III

← Needs to shift left
shifts $2BaO_2(s) \rightleftharpoons 2BaO(s) + O_2(g)$

$$\Delta H^\circ = 162 \text{ kJ/mol}_\text{rxn} \text{ endothermic}$$

A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ *R*
(B) Removing a small amount of $BaO(s)$ *X*
(C) Adding He gas to the vessel *X*
(D) Lowering the temperature *L*

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