

Physical Science: Chapter 11  
Section 2

**Acceleration**

# Acceleration Study Guide Section 2 Physical Science

**James L. Meriam, J. M. Henderson**



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web this experimental approach required the introduction of a new unit for amount of substances the mole which remains indispensable in modern chemical science the mole is an amount unit similar to familiar units like pair dozen gross etc it provides a specific measure of the number of atoms or molecules in a bulk sample of matter

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d 0 3 mol so<sub>4</sub> 2 c 3 calculate mass of 1 mol electron c 4 how many moles of electron will weight of 1 kg c 5 how many moles  
of proton will weight 200 g

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web the amount in moles of pb 3 o 4 to 3 significant figures iii the amount in moles of lead produced iv mass of lead produced  
to 3 significant figures b 35 61 g of tin reacts with exactly 42 60 g of chlorine cl 2 to form 78 21 g of tin iv chloride sncl 4 i  
calculate the number of moles of tin chlorine and tin chloride a r values

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substance is atomic or avogadro s number of molecules if the substance is molecular 1 mole of carbon atoms 6 022 10<sup>23</sup>  
atoms of carbon 1 mole of sodium atom 6 022 10<sup>23</sup> atoms of sodium

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number of particles atoms molecules or ions it contains it is based on avogadro s number which states that one mole of any  
substance contains 6 022 x 10<sup>23</sup> particles 2