

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HBr~~ **hydrobromic acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~HClO₄~~ **perchloric acid**
- ~~Sodium hydroxide~~ **NaOH**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~barium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **NaOH**

~~HClO₄ is a STRONG acid & it won't completely dissociate in water.~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HBr~~

~~HNO₃~~

~~Mg(OH)₂~~

NaOH

NaOH

~~NaOH~~

NaOH

3. What is the formula for hydronium? **H₃O⁺**

4. $\text{H}_2\text{C}_2\text{O}_4$ will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{H}_2\text{C}_2\text{O}_4 + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{C}_2\text{O}_4^+ + \text{HC}_2\text{O}_4^-$
- Would you expect a large or small equilibrium constant for this reaction? **small, because $\text{H}_2\text{C}_2\text{O}_4$ is a weak acid.**
- Would you define $\text{H}_2\text{C}_2\text{O}_4$ as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is $\text{H}_2\text{C}_2\text{O}_4$ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of $\text{H}_2\text{C}_2\text{O}_4$: **HC_2O_4^-**
- Identify $\text{H}_2\text{C}_2\text{O}_4$ as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HSO_4^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{HSO}_4^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{SO}_4 + \text{SO}_4^{2-}$
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO_4^- is a weak acid.**
- Would you define HSO_4^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HSO_4^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO_4^- : **SO_4^{2-}**
- Identify HSO_4^- as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and can take a proton) and it has H's to give (and can lose one).**

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