

Chemical Kinetics

1). Molecular hydrogen can be made from methane gas by the reaction below. How is the rate of disappearance of CH_4 related to the rate of appearance of H_2 ?



- (A) $\text{rate of disappearance of } \text{CH}_4 = \text{rate of appearance of } \text{H}_2$
- (B) $\text{rate of disappearance of } \text{CH}_4 = 3 \times \text{rate of appearance of } \text{H}_2$
- (C) $\text{rate of disappearance of } \text{CH}_4 = \frac{1}{3} \times \text{rate of appearance of } \text{H}_2$
- (D) none of these

Answer: B

2). For the reaction shown below, what is the relationship between the rate of formation of Br_2 and the rate of reaction of Br^- over the same time period?



- (A) rate of formation of $\text{Br}_2 = 1/2 \times$ rate of reaction of Br^-
- (B) rate of formation of $\text{Br}_2 = 2/4 =$ rate of reaction of Br^-
- (C) rate of formation of $\text{Br}_2 = 2 \times$ rate of reaction of Br^-
- (D) rate of formation of $\text{Br}_2 = 2 \times$ rate of reaction of Br^-

Answer: A

3). The reaction that occurs in a Breathalyzer, a device used to determine the alcohol level in a person's bloodstream, is given below. If the rate of appearance of $\text{Cr}_2(\text{SO}_4)_3$ is 3.20 mol/min at a particular moment, what is the rate of disappearance of $\text{C}_2\text{H}_5\text{OH}$ at that moment?



- (A) 0.400 mol/min
- (B) 0.600 mol/min
- (C) 1.66 mol/min
- (D) 3.72 mol/min

Answer: C

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David Wilson

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