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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$m_{mix} = \frac{m_1 \cdot V_1 + m_2 \cdot V_2}{V_1 + V_2}$$

$$1.5 \times 480 + 1.2 \times 520 = 480 + 520 = 1000$$

$$1.344 \text{ M}$$
 with increase in temperature which of these changes a molality b fraction of solute present in water

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