

## ACS Practice Test 1

### Acids & Bases:

1. All are examples of Lewis acid–base reactions *except*

- (A)  $\text{Cu}^{2+}(\text{aq}) + 4\text{NH}_3(\text{aq}) \rightleftharpoons [\text{Cu}(\text{NH}_3)_4]^{2+}(\text{aq})$   
 (B)  $\text{HCl}_{(\text{g})} + \text{NH}_{3(\text{g})} \rightarrow \text{NH}_4\text{Cl}_{(\text{s})}$   
 (C)  $\text{H}^{+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{H}_2\text{O}_{(\text{l})}$   
 (D)  $2\text{Na}_{(\text{s})} + \text{Cl}_{2(\text{g})} \rightarrow 2\text{NaCl}_{(\text{s})}$

2. According to the Lewis definition, an acid is a species

- (A) having a hydrogen ion. (B) donating a pair of electrons.  
 (C) accepting a pair of electrons. (D) accepting a hydrogen ion.

3. All are potential Lewis bases *except*

- (A)  $\text{NH}_3$  (B)  $\text{H}_2\text{O}$  (C)  $\text{CH}_4$  (D)  $\text{CN}^{-}$

4. Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?

- (A)  $\text{Al}^{3+}$  (B)  $\text{Ba}^{2+}$  (C)  $\text{K}^{+}$  (D)  $\text{Zn}^{2+}$  (E)  $\text{Ag}^{+}$

5. Which element in Group 5A forms the most basic oxide?

- (A) N (B) P (C) As (D) Sb (E) Bi

### Atomic Structure:

6. In what respect does an atom of magnesium differ from a magnesium ion ( $\text{Mg}^{2+}$ )?

- (A) The ion has a more stable electronic arrangement than the atom.  
 (B) The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.  
 (C) The ion has two more protons than the atom.  
 (D) The ion will react more readily with free  $\text{Cl}_2$ .  
 (E) The ion has two more planetary electrons than the atom.

7. A certain atom has the symbol  ${}^{148}_{73}\text{Pc}$ . What does this symbol tell about an atom of the element?

- (A) It has 148 electrons. (B) It has a valence of 4.  
 (C) It has 221 protons in the nucleus. (D) It has 75 neutrons in the nucleus.  
 (E) Its atomic mass is 73 u.

8. The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is

- (A) 16 (B) 31 (C) 32 (D) 48

9. The atomic mass of an element is 32.07 u and its atomic number is 16. The element forms a simple ion. The ion will most likely have a charge of

- (A) 1– (B) 2– (C) 3+ (D) 1+

# Acs Practice Test 2 Answer Key

**Pat Floyd, Patricia A. Floyd, Sandra E.  
Mimms, Caroline Yelding-Howard**

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**Field Manuals** United States. War Department,1948      Test Preparation Guide for Course 2 Martha Humbard,1995

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**Personal Health** Pat Floyd,Patricia A. Floyd,Sandra E. Mimms,Caroline Yelding-Howard,2003 This book emphasizes the individual's personal responsibility for wellness by presenting general core and current health information to guide decision making The focus of the text is to present health topics including cultural racial ethnic and gender diversity identify risk factors and give students useful and sensible suggestions to reduce their risk for preventable diseases and conditions in order to achieve optimal levels of wellness for themselves friends and family Filled with practical advice stimulating discussion questions and self assessment exercises this text takes an applied approach and empowers students to take control of their own health Concentrated emphasis on health issues as they relate to various cultural differences make this text one of the most up to date resources for health related culture gender ethnic and age issues

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