

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. acid rain | 6. buffer solution |
| 7. ionization | 8. weak solution | 9. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

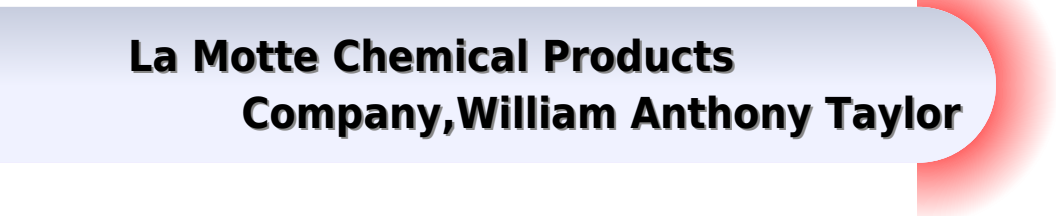
Write the equation and/or the chemical formula(s) of the species involved in the reaction described in each problem. Be sure to include the correct state symbols.

1. Write the equation for the autoionization of water.
2. Write the equation for the autoionization of water which produces an acidic solution.
3. Write the equation for the autoionization of water which produces a basic solution.
4. Write the equation for the autoionization of water which produces a neutral solution.
5. Write the equation for the autoionization of water which produces a basic solution.
6. Write the equation for the autoionization of water which produces a basic solution.
7. Write the equation for the autoionization of water which produces a basic solution.
8. Write the equation for the autoionization of water which produces a basic solution.
9. Write the equation for the autoionization of water which produces a basic solution.
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11. Write the equation for the autoionization of water which produces a basic solution.
12. Write the equation for the autoionization of water which produces a basic solution.
13. Write the equation for the autoionization of water which produces a basic solution.
14. Write the equation for the autoionization of water which produces a basic solution.
15. Write the equation for the autoionization of water which produces a basic solution.
16. Write the equation for the autoionization of water which produces a basic solution.
17. Write the equation for the autoionization of water which produces a basic solution.
18. Write the equation for the autoionization of water which produces a basic solution.
19. Write the equation for the autoionization of water which produces a basic solution.
20. Write the equation for the autoionization of water which produces a basic solution.

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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19. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
20. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

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**La Motte Chemical Products
Company, William Anthony Taylor**



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václav eusebius z valdštejna 24 september 1583 in hermanitz an der elbe bidschower kreis 25 februar 1634 in eger elbogener kreis war ein böhmischer feldherr generalissimus und ständepolitiker zudem war er herzog von [albrecht von wallenstein bohemian war assassination legacy](#) - Nov 06 2022

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