

## CHEM 12: Hess's Law Worksheet

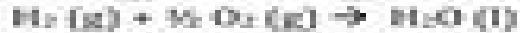
1. Calculate the  $\Delta H$  for the reaction:  $\text{C}_2\text{H}_6(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g})$ , from the following data:



$$\Delta H = -1411 \text{ kJ / mole}$$



$$\Delta H = -1560 \text{ kJ / mole}$$



$$\Delta H = -285.8 \text{ kJ / mole}$$

2. Calculate the  $\Delta H$  for the reaction:



from the following data:



$$\Delta H = -180.5 \text{ kJ / mole}$$



$$\Delta H = -91.8 \text{ kJ / mole}$$



$$\Delta H = -483.6 \text{ kJ / mole}$$

3. Find the  $\Delta H_f^\circ$  (heat of formation) for acetic acid,  $\text{HC}_2\text{H}_3\text{O}_2$ , using the following thermochemical data:



$$\Delta H = -875 \text{ kJ / mole}$$



$$\Delta H = -394.51 \text{ kJ / mole}$$



$$\Delta H = -285.8 \text{ kJ / mole}$$

Pure Factor C(s) is in the form of graphite.

Acetic acid,  $\text{HC}_2\text{H}_3\text{O}_2$ , actually has the structure:  $\text{CH}_3\text{COOH}$



4. Calculate the  $\Delta H$  for the reaction:  $\text{CH}_4(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{HCN}(\text{g}) + 3\text{H}_2(\text{g})$ , from the following data:



$$\Delta H = -91.8 \text{ kJ / mole}$$



$$\Delta H = -74.9 \text{ kJ / mole}$$



$$\Delta H = +270.3 \text{ kJ / mole}$$

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**Disha Experts**

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