

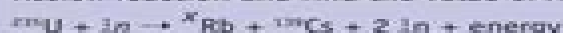
Representing and interpreting nuclear reactions

We can represent nuclear reactions in the form of equations such as ${}^{14}_6\text{C} \rightarrow {}^{14}_7\text{N} + {}^0_{-1}\beta$

This equation represents the decay of a carbon-14 nucleus leaving a daughter nitrogen nucleus and emitting a beta particle in the process. Notice that the equation is 'balanced' in the sense that the total mass number and atomic number is the same on both sides of the arrow. This must always be the case. We know that energy is conserved in the reaction, so the energy released when a carbon-14 nucleus decays to a nitrogen nucleus is emitted in the form of a beta particle.

Worked example 8.9

Complete the following equation which represents a nuclear fission reaction and find the value of X.



The equation is balanced on both sides so the total mass number on the left side must be the same as the total mass number on the right side

$$235 + 1 = X + 141 + 2$$

$$236 = X + 141$$

$$236 - 141 = X$$

$$= 95$$

We can combine nuclear reaction equations and data about the mass of the elements on both sides of the equation to find out how much energy is released per fission.

Worked example 8.10

Calculate the energy released in the following fission reaction. Give your answer in MeV.



The data you need are:

mass of U-235 is 235.0439 u

mass of Ba-141 is 140.9144 u

mass of Kr-92 is 91.9262 u

mass of ${}^1_0\text{n}$ is 1.008 665 u.

Consider mass on each side of the equation

$$235.0439 \text{ u} + 1.008 66 \text{ u} \rightarrow 140.9144 \text{ u} + 91.9262 \text{ u} + (3 \times 1.008 665 \text{ u})$$

$$236.05256 \text{ u} \rightarrow 235.866595 \text{ u}$$

$$\Delta m = (236.05256 - 235.866595) \text{ u} \\ = 0.185 965 \text{ u}$$

$$\text{Energy released} = 0.185 96 \times 931.5 \text{ MeV} \\ = 173 \text{ MeV}$$

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Erin McCarthy



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